

1 Nuclear Chemistry

Radioactive Cat. Has 18 Half-Lives.

Types of Reactions:

Chemical Reactions:

- involve valence electrons
- elements are rearranged to form new compounds
- Law of Conservation of Mass obeyed
- relatively little energy can be absorbed or released

Nuclear Reactions:

- changes occur in the nucleus; involve protons, neutrons, and electrons
- Transmutation occurs (one element turns into another)
- Law of Conservation of Mass is not always obeyed; mass can be changed into energy
- lots of energy can be produced

Roentgen discovered x-rays in 1895 as a result of exposing photographic plates to the rays emitted from a cathode ray tube. He found that x-rays could travel through some substances.

Becquerel in around 1900 studied phosphorescence (glowing) in uranium ore when he discovered that the ore also exposed photographic plates.

The Curies further experimented with uranium ore. They isolated uranium as the substance with the unique properties. They were the first to call uranium "radioactive" - meaning a substance that gave off rays of energy from the nucleus.

Radioisotopes are isotopes that have unstable nuclei; they emit radiation to become more stable.

The radiation emitted can be one of several types. The three most common types of radiation are:

- alpha, α
- beta, β
- gamma, γ

Alpha decay occurs when an atom loses an alpha particle (2 protons + 2 neutrons), aka a helium nucleus. The loss of two protons decreases the atomic number by 2 and decreases the mass number by 4. Since the number of protons changes, a new element is formed. Alpha radiation isn't powerful (compared to other types of radioactive decay) and can be stopped by paper alone.

Beta decay occurs when an atom loses an electron from the nucleus. It is the result of a neutron converting to a proton and electron and the high energy electron is ejected. The new proton increases the atomic number by 1, which results in a new element. Note that the mass number is unchanged.

Gamma radiation occurs when high energy photons (electromagnetic energy) are ejected from the nucleus. Usually accompanies other nuclear emissions and represents the energy lost when the nucleus becomes more stable. Does not change the atomic number or mass and it is difficult to stop; it requires several cm of lead or several feet of concrete to block.

Relative Strengths of Radiation:

- Alpha particles can be stopped by paper
- Beta particles can be stopped by a few mm of aluminum

- Gamma particles can be stopped by a few cm of lead or several feet of concrete

Nuclear Mass Defect - The difference in the calculated mass of an atom (based on C-12) and the measured mass. Measured mass is always smaller than expected. Lost mass has been converted to energy.

Nuclear Binding Energy - the energy that comes from the lost mass. Holds the nucleus together. Lots of energy produced, according to Einstein's famous equation, $E = mc^2$

Strong Nuclear Force - a force that keeps the protons from repelling each other in the nucleus. Acts over a short distance and overcomes electrostatic repulsion (like charges repel). Strong nuclear force struggles in very large elements.

Stability of Isotopes:

- Nuclei that are not stable tend to undergo radioactive decay, a process that emits radiation and particles from the nucleus
- Spontaneous decay is the result of instability in the nucleus of an atom.

What determines nuclear stability?

- Neutron to proton ratio. If this ratio is close to 1:1 in a low mass (small) element, the isotope is likely to be stable.
- Neutron to proton ratio. For larger elements, the stable ratio is 1.5: 1-1.5 neutrons to protons

In a graph of protons vs neutrons, the area of the graph within which all stable nuclei are found is called the band of stability.

- Nuclei with a high neutron to proton ratio will undergo beta decay.
- Nuclei with low ratio will undergo positron emission or electron capture.
- Nuclei with high atomic masses tend to undergo alpha decay.

Electron Capture:

- The nucleus captures an electron from the emission cloud
- The electron combines with a proton to convert into a neutron
- The mass number remains the same and the atomic number decreases since a proton was converted

Positron Emission:

- The radioactive decay process in which a positron is emitted from the nucleus.
- A positron is a particle with the same mass as an electron but opposite in charge.
- The number of protons decreases but the mass number remains the same.

Artificial Nuclear Changes Neutron Bombardment

- A neutron enters the nucleus at a high speed and is captured; usually creates a radioactive isotope
- Creates synthetic isotopes useful in medicine and research
- Increases the mass number by one, but the atomic number remains the same.

A radioactive decay series is a series of nuclear reactions that begins with an unstable nucleus and ends with the formation of a stable nucleus.

- Nuclei that are not stable tend to undergo radioactive decay, a process that emits radiation (gamma radiation) and particles from the nucleus
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Most elements have some radioactive isotopes (radioisotopes) that occur naturally, but the percentage of radioactive isotopes is so low that we do not consider that element radioactive.

Other elements, such as radon, have a high percentage of radioisotopes, so we consider that element, in general, to be radioactive.

Radioactive isotopes can be created from stable ones by neutron bombardment, which causes instability in the nucleus by disturbing the 1:1 neutron:proton ratio.

All elements heavier than bismuth are naturally radioactive due to their instability. They tend to spontaneously undergo alpha decay to bring them closer to a stable nucleus.

All elements heavier than uranium, atomic number 92, are artificially created and are always radioactive. These are the transuranium elements. They are usually formed by bombarding a heavy atom with neutrons or alpha particles. Scientists have created elements up to about 118 protons.

Half life is the amount of time it takes for one half of a radioactive element to change from the parent element to the daughter element (from radioactive element to product element).

Use for formulas

$$m_t = m_0(1/2)^n$$
$$n = \frac{t}{T}$$

where m_t is the final mass, m_0 is the original mass, n is the number of half lives, t is the elapsed time and T is the time for one half life.

You can also use the following formula

$$\ln(m_t) = -(0.693/T)t + \ln(m_0)$$

- Fossils are dated by knowing the ratio of parent element to daughter element
- Two things must be known - length of half life and ratio of parent:daughter
- The carbon-14 isotope is used for wood, bones, tissue, anything that used to be alive
- All living things contain carbon, and part of the carbon is C-14
- After something dies, it stops exchanging carbon with the atmosphere and the C-14 begins to decay
- Because we know the half-life of C-14 to be 5730 years, we can measure the amount of C-14 left and calculate how long the organism has been dead.

Nuclear Fission:

- Fission occurs when a neutron hits a large, unstable nucleus and causes it to break into two smaller nuclei.
- Extra neutrons are emitted from this reaction, and those neutrons go on to hit other nearby atoms, causing them to go through fission.
- A chain reaction occurs if there are enough atoms present.

Critical mass is the amount of a substance that must be present in order for a chain reaction to occur.

If not enough mass is present, the neutrons will escape the sample without hitting other atoms, and there will be no chain reaction. This is called subcritical mass.

In addition to neutrons being released, large amounts of energy are also released. Nuclear fission is highly exothermic. In nuclear power plants, the energy and chain reaction is controlled and slow. In an atom bomb, the fission is uncontrolled and very fast.

Parts of a Nuclear Reactor in a Power Plant

- Fuel rods - usually uranium-235 or plutonium, a fissionable isotope
- Control rods - surround the fuel rods and absorb some of the neutrons to control the chain rxn
- Graphite Moderator - slows the neutrons so they can enter the fuel atoms at the most effective rate
- Water heated to steam - same as fossil fuel plant. The exothermic reaction heats the water to steam, which turns a turbine
- Concrete Containment Building - dome shaped building around the nuclear reactor to keep radioactivity inside

Nuclear Fusion

- Fusion is the process by which small nuclei fuse together to form a heavier nucleus
- This is the main process that goes on in our sun and other stars. It is responsible for the light and energy given off - highly exothermic for small atoms
- The primary reaction is two hydrogen isotopes building helium

Fusion produces lots of energy, and its products are generally not radioactive. It is appealing as an energy source. However, to get fusion started, lots of energy is needed to get two repelling hydrogen nuclei to move together. A thermonuclear reactor is required. Fusion isn't a feasible energy source for now.

Uses of Radioactivity

- Radiation therapy for cancer - radiation targets cancerous cells
- Radiotracers for medical diagnosis and research
- Nuclear Power
- Food preservation - it doesn't make your food radioactive, it kills bacteria and fungus
- Smoke Detectors

Properties of Radioactive Elements

- Radioactive elements will expose photographic film
- Ionize other atoms, causing them to become charged
- Can cause some other substances to fluoresce, or glow in the dark
- Prolonged exposure can cause damage to cells

2 types of radiation damage:

Somatic

- affects nonreproductive body tissue
- Only affects organism during its lifetime
- Causes burns, cancer

Genetic

- Damages sex cells
- Affects chromosomes and therefore offspring
- Causes mutations, lifelong disease, and deformities

Chapter Problems

1. Explain the relationship between an atom's neutron-to-proton ratio and its stability.
2. Write a balanced nuclear equation for the alpha decay of americium-241.
3. Technetium-104 has a half-life of 18.0 minutes. How much of a 165.0 sample remains after 90.0 minutes?
4. A sample initially contains 150.0 mg of radon-222. After 11.4 days, the sample contains 18.7 mg of radon-222. Calculate the half-life.
5. Discuss how the amount of fissionable material present affects the likelihood of a chain reaction.
6. Describe the penetration power of alpha, beta, and gamma radiation.